**Naming Rules Card**

*The following are all the chemical compounds naming rules that we will be using in this class (there are more of course).* If you turn in the note card to get it **laminated** you will be able to use it on **ANY** **quiz or test for the rest of the year!**  Any non-laminated cards will not be able to be used.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| Monatomic Cations | | | | Molecular Compounds |
| Element | **Formula** | **Element** | **Formula** | 1= mono |
| Aluminum | Al3+ | Lead(II)/Plumbous | Pb2+ | 2 = di |
| Barium | Ba2+ | Lead(IV)/Plumbic | Pb4+ | 3 = tri |
| Cadmium | Cd2+ | Maganese(II)/Manganous | Mn2+ | 4 = tetra |
| Calcium | Ca2+ | Maganese(III)/Manganic | Mn3+ | 5 = penta |
| Cesium | Cs+ | Magnesium | Mg2+ | 6 = hexa |
| Chromium(II)/Chromous | Cr2+ | Mercury(I)/Mercurous | Hg22+ | 7 = hepta |
| Chromium(III)/Chromic | Cr3+ | Mercury(II)/Mercuric | Hg2+ | 8 = octo |
| Cobalt(II)/Cobaltous | Co2+ | Potassium | K+ | 9 = nano |
| Cobalt(III)/Cobaltic | Co3+ | Rubidium | Rb+ | 10 = deca |
| Copper(I)/Cuprous | Cu+ | Silver | Ag+ | **\*\*Drop Double Vowels** |
| Copper(II)/Cupric | Cu2+ | Sodium | Na+ |  |
| Gold | Au+ | Strontium | Sr2+ |  |
| Hydrogen | H­+ | Tin(II)/Stannous | Sn2+ |  |
| Iron(II)/Ferrous | Fe2+ | Tin(IV)/Stannic | Sn4+ |  |
| Iron(III)/Ferric | Fe3+ | Zinc | Zn2+ |  |
| Lithium | Li+ |  |  |  |
|  | | | |  |
| Polyatomic Cations | | | |  |
| Ammonium | NH4+ |  |  |  |
| Hydronium | H3O+ |  |  |  |

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| **Monatomic Anions** | | **Polyatomic Anions** | | | |
| **Element** | **Formula** | **Name** | **Formula** | **Name** | **Formula** |
| Hydride\* | H- | Acetate (ethanoate) | CH3CO2-or C2H3O2- | Hypochlorite | ClO- |
| Fluoride\* | F- | Bromate | BrO3- | Iodate | IO3- |
| Chloride\* | Cl- | Carbonate | CO­32- | Nitrate | NO3- |
| Bromide\* | Br- | Chlorate | ClO3- | Nitrite | NO2- |
| Iodide\* | I- | Chlorite | ClO2- | Oxalate (ethanedioate) | C2O42- |
| Oxide\* | O2- | Chromate | CrO42- | Perchlorate | ClO4- |
| Sulfide | S2- | Cyanide | CN- | Permanganate | MnO4- |
| Nitride\* | N3- | Dichromate | Cr2O72- | Phosphate | PO43- |
| \*Exists as diatomic molecules | | Dihydrogen Phosphate | H2PO4- | Phosphite | PO33- |
| **Naming Acids** | | Hydrogen Carbonate (bicarbonate) | HCO3- | Silicate | SiO32- |
| **Monatomic Anion** | | Hydrogen Phosphate (biphosphate) | HPO42- | Sulfate | SO42- |
| \*Name starts with hydro- | | Hydrogen Sulfate (bisulfate) | HSO4- | Sulfite | SO32- |
| Ion has –ic at the end and add “acid” | | Hydroxide | OH- | Thiocyanate | SCN- |
| **Polyatomic Anion** | |  |  | Thiosulfate | S2O32- |
| \*No hydro in name | |  |  |  |  |
| If ion ends in –ite change it to –ous and add “acid” | |  |  |  |  |
| If ion ends in –ate change it to –ic and add acid | |  |  |  |  |
| Remember: “I –ate something –icky” | |  |  |  |  |